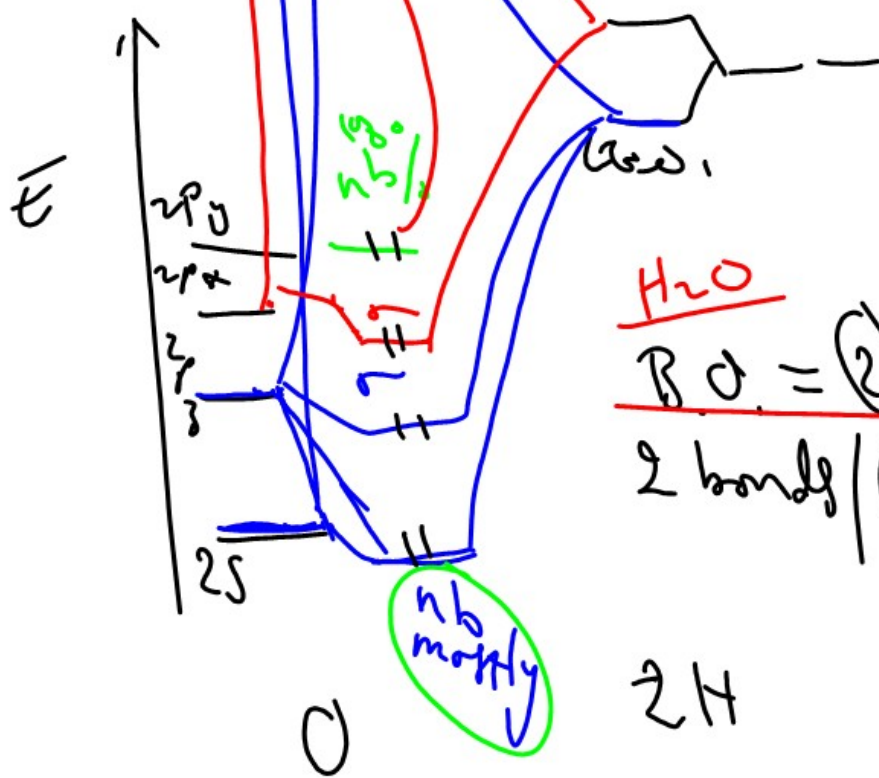
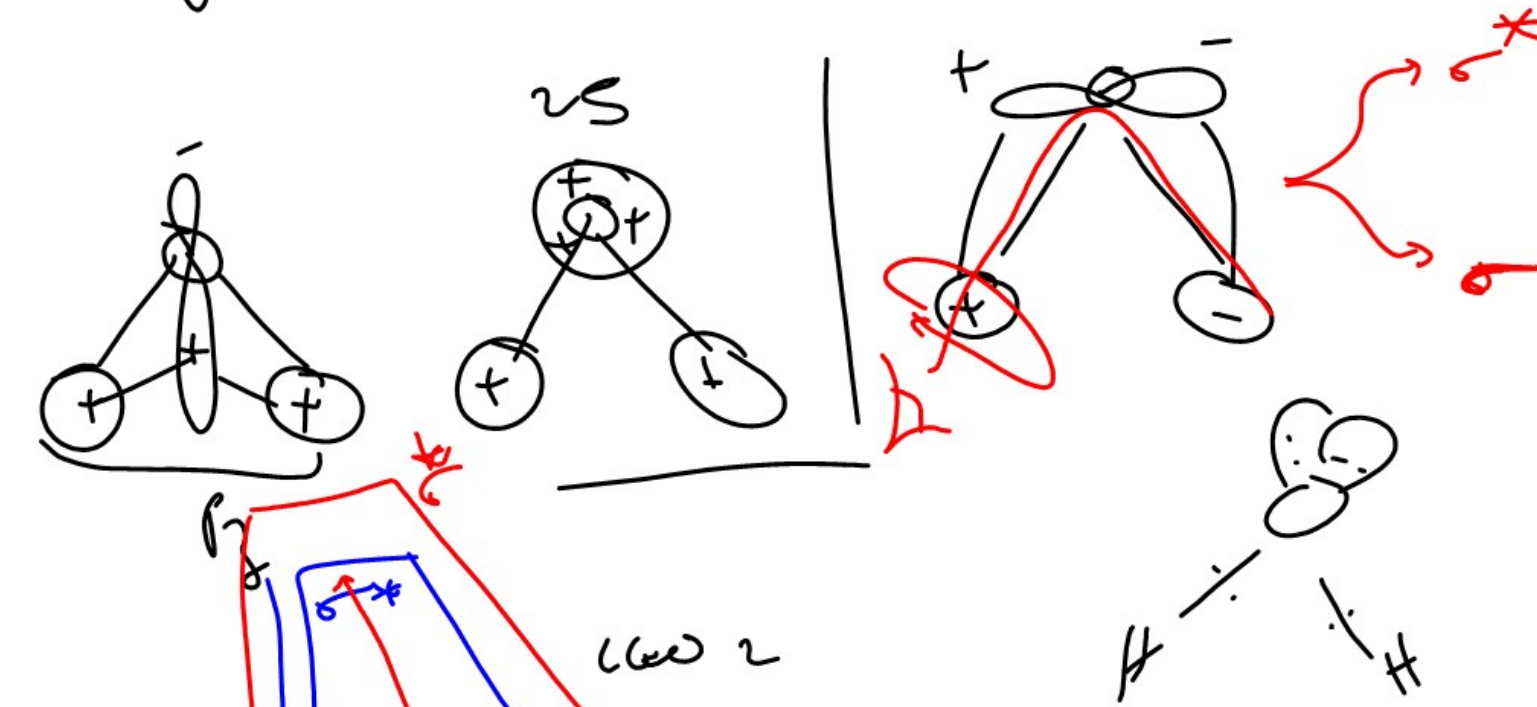
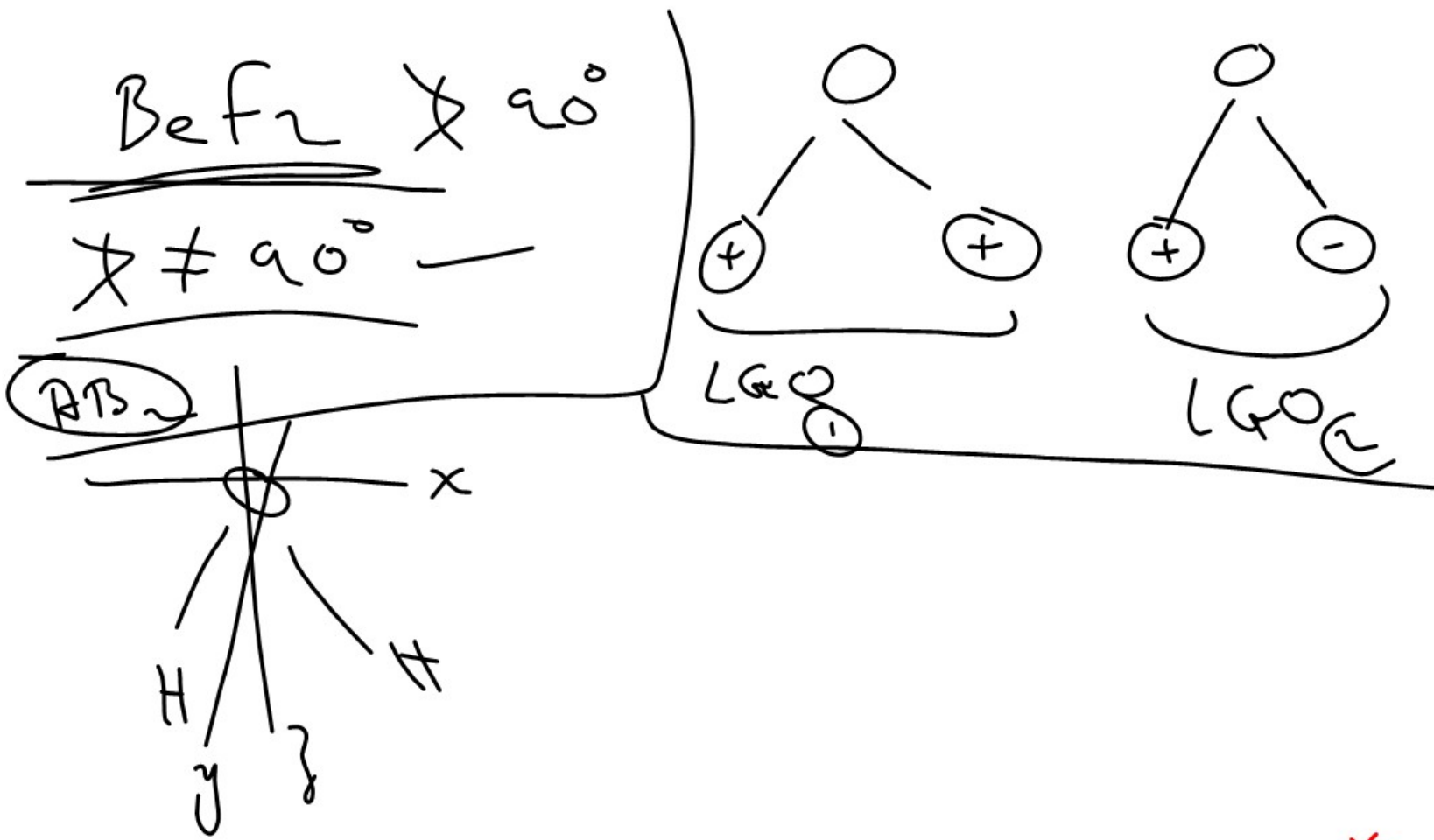
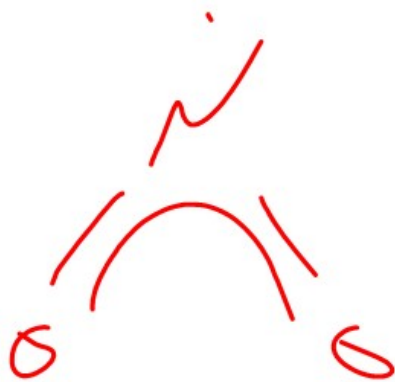
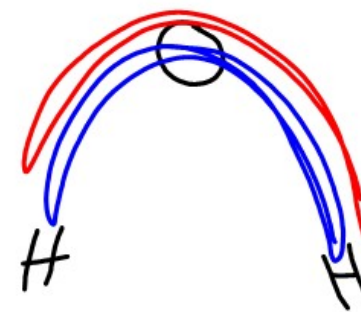




H₂O
B.O. = 2
2 bonds / (3C - 2e bond)

H₂O*
B.O. = 3.5

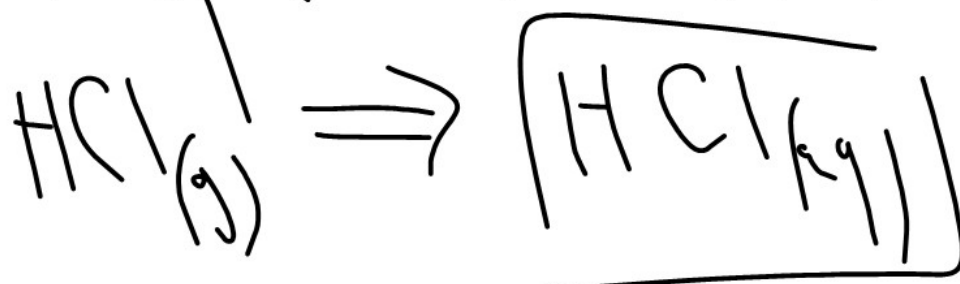


Acid-Base (Chap^{tr}. 6)

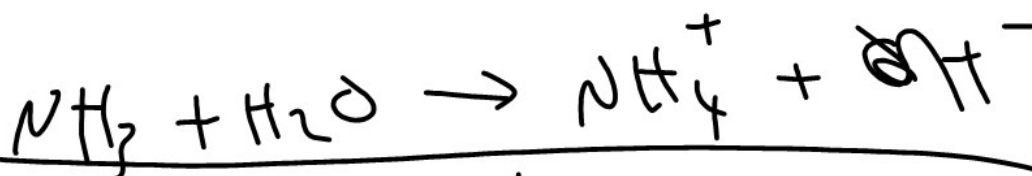
Based on M.O.T.

Arrhenius Def.

acid: subst. that increases $[H_3O^+]$ in aqueous media



base: subst. increases $[OH^-]$ in aq. media



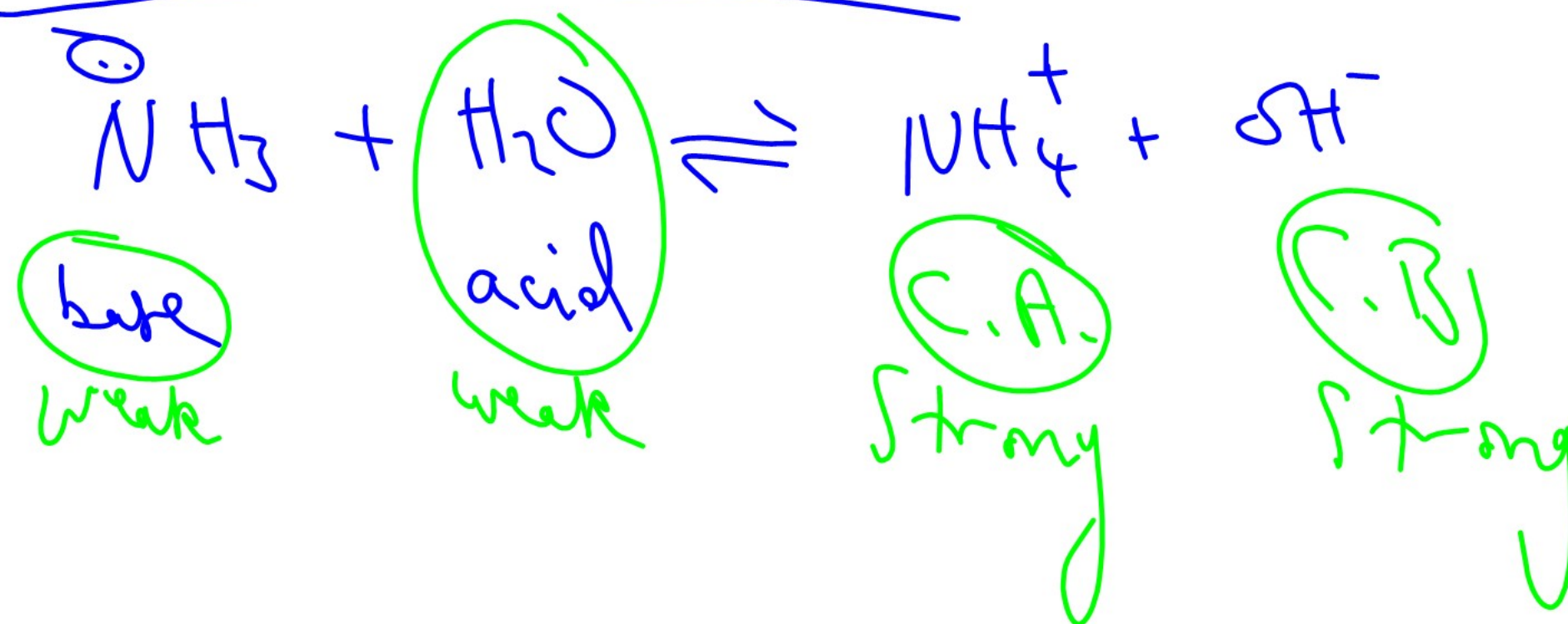
Arrhenius Def. is narrow

Bronsted-Lowry {aqueous}

acid: Proton donor

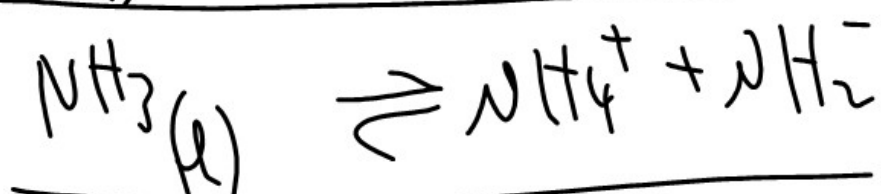
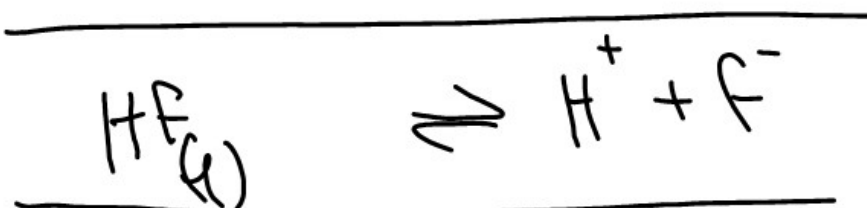


base: Proton acceptor



③ Solvent System Def.

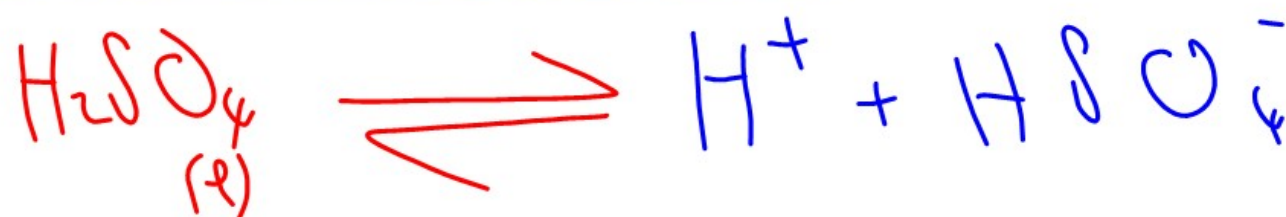
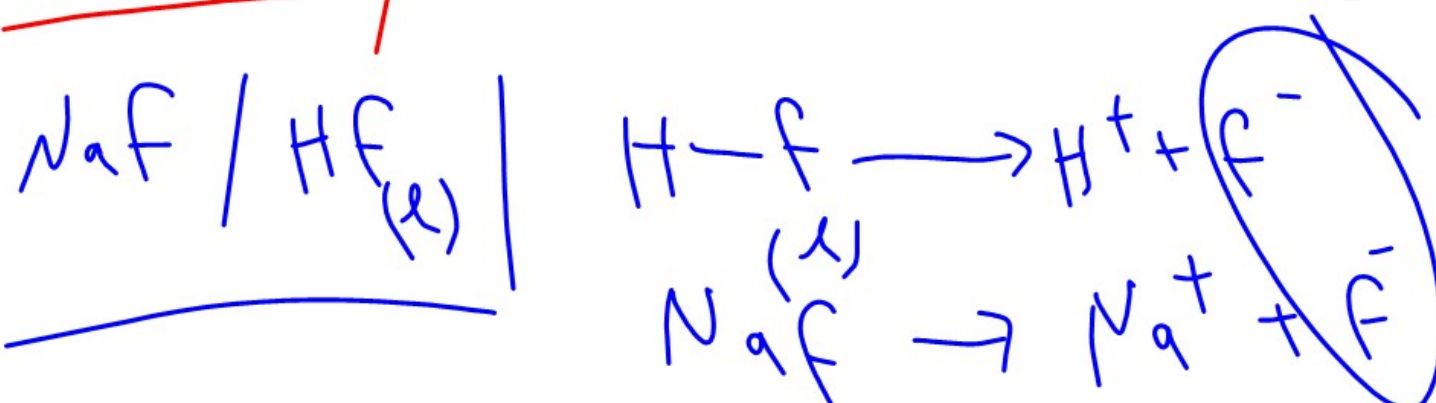
autoionizing

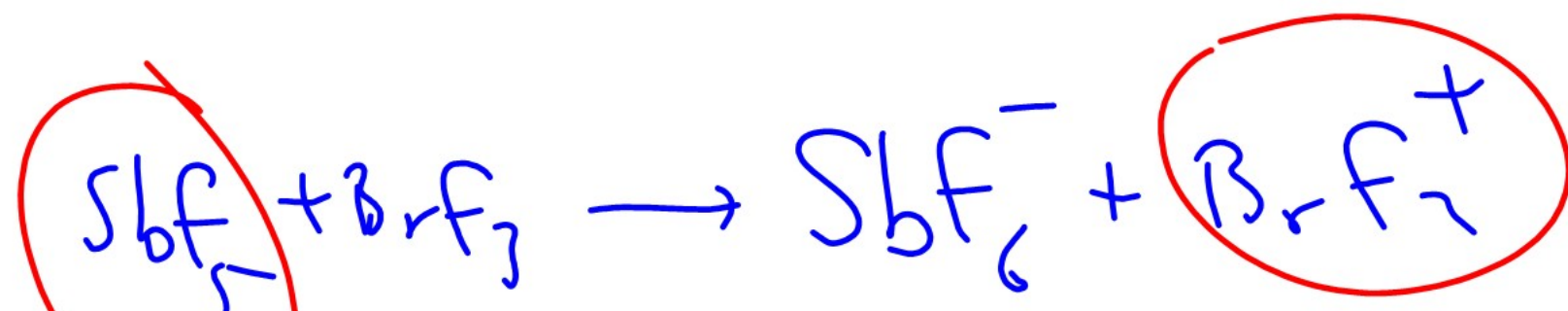


acid: subst. that increases (cation)
base: subst. ' ' (anion)

H_2SO_4 in water $\Rightarrow$ acidic

NaNH_2 / inside $\text{NH}_3_{(l)}$





Solvent $\rightarrow$ autoioniz. $\Rightarrow$ Protic
 $\rightarrow$ aprotic
 ~~$\rightarrow$ nonautoioniz.~~